**List of Table Captions**

Table 1. Conditions under which experiments of the reaction kinetics of [AuCl4]- with formic acid have beeen carried out.

Table 2. Influence of the reductant concentration on the observed rate constant (*k*obs) for the reaction of [AuCl4]- with HCOOH in the solution with pH = 2.9. Conditions: *C*0,Au(III) = 0.15 mM; temperature (50±0.1) oC.

Table 3. The observed rate constant for the reaction of [AuCl4]- with HCOOH at different initial concentration of [AuCl4]- in solution with pH = 2.9. The other experimental conditions: 100-fold excess of reductant concentration; *I* = 0.05 M NaClO4; temperature (50±0.1) oC.

Table 4. The values of the rate constant (*k*obs) for the reaction of [AuCl4]- with HCOOH at different ionic strength. Experimental conditions: pH = 2.9; temperature (50±0.1) oC; *C*0,Au(III) = 0.15 mM; *C*0,HCOOH = 1.5 mM.

Table 5. The values of the rate constant (*k*obs) for the reaction of [AuCl4]- with HCOOH at different temperatures. Experimental conditions: *C*0,Au(III) = 0.15 mM; *C*0,CHOOH = 1.5 mM; pH = 2.9; *I* = 0.05 M.

Table 6. The values of the rate constant for the reaction of [AuCl4]- with HCOOH at different Cl- concentration. Experimental conditions: *C*0,Au(III) = 0.15 mM; *C*0,HCOOH = 0.75 mM; pH = 2.9; *I* = 0.05 M, temperature (50±0.1) oC.